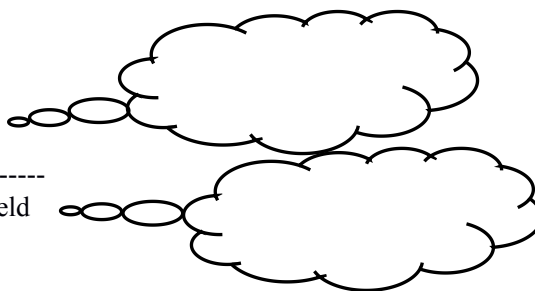


## 6.5 Percent Yield (S) Notes

$$\% \text{ yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}}$$



**Percent yield** shows how much you got out of a reaction compared to how much you expected to get.

Examples:

After having done some stoichiometry, Mr. Newman calculated that he should get 5.00 g of table salt (NaCl). Instead, in the lab, he only got 4.37 g. Find the percent yield.

Car manufacturers know that there's typically a 98% yield of the steel after a certain procedure to make it stronger. How much steel will they get if the theoretical yield is 155 kg of steel?

You are working in lab and your lab partner informs you that the % yield of this reaction was an abysmal 35%. You weigh the mass of the product and find that it is 0.488 g. You haven't done the stoichiometry yet because you forgot to do your HW, but what should the mass of your product be (in the stoichiometry problem) based on your lab partner's comment?