

6.4 Limiting Reactants (S) Prequiz

1. The substance that restricts the participation of other reactants in a chemical reactions is known as the:

- a) Limiting reactant
- b) Limiting product
- c) Excess reactant
- d) Excess product

2. To determine the limiting reactant in a chemical reaction, one must know the:

- a) Available amount of one of the reactants
- b) Amount of product formed
- c) Available amount of each reactant
- d) Speed of the reaction

For the next few questions, use this chemical equation: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

3. Suppose you have **3.0** moles of hydrogen gas and **3.0** moles of oxygen gas. What is the limiting reactant?

- a) H_2
- b) O_2
- c) H_2O
- d) There isn't one in this situation.

4. Suppose you have **2.0** moles of hydrogen gas and **3.0** moles of oxygen gas. What is the limiting reactant?

- a) H_2
- b) O_2
- c) H_2O
- d) There isn't one in this situation.

5. Suppose you have **3.0** moles of hydrogen gas and **2.0** moles of oxygen gas. What is the limiting reactant?

- a) H_2
- b) O_2
- c) H_2O
- d) There isn't one in this situation.

6. Suppose you have 3.0 **grams** of hydrogen gas and 3.0 **grams** of oxygen gas. What is the limiting reactant?

- a) H_2
- b) O_2
- c) H_2O
- d) There isn't one in this situation.