4.2 Naming & Formulas of Ionic & Covalent Compounds (E) Notes

Intro: Ions

Why? “Octet Rule” & Valence Electrons

Ionic Compounds

**Naming**

If there are 2:

Cation Anion

If there are 3:

\*Look for the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

(You can use your sheet!)

KNO3

NH4F

**Formula**

Rule: must have same # of \_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_

What does the formula mean?

More Practice (on pg \_\_\_\_\_\_\_\_\_\_)

Special Rule:

Roman Numberals and the d-block

Covalent Compounds

**Naming**

\*\*First Second

**Formula**

|  |  |
| --- | --- |
| Prefix | # of Atoms |
|  | 1\*\* |
|  | 2 |
|  | 3 |
|  | 4 |
|  | 5 |

\*\*Special rule for the first atom:

Why do we *not* use positives & negatives?

More Practice (on page \_\_\_\_\_\_\_\_\_\_)

Question: How do I know whether something is Ionic or Covalent?