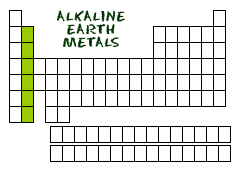


**Alkali Metals** are in group 1, but they do not include Hydrogen!!

(So make sure people do not include Hydrogen when marking their Periodic table!

Properties: (1) They are very, very reactive (they explode when they touch water!)

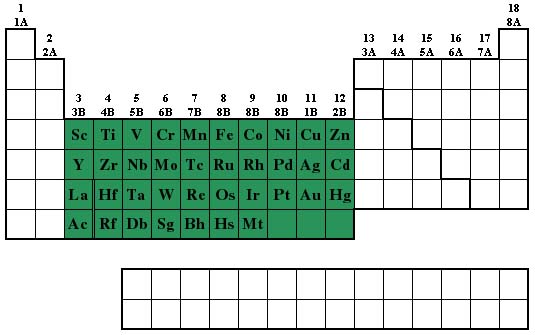
1. They are a metal (duh).



**Alkaline-Earth Metals** are in group 2.

Properties: (1) They are very reactive.

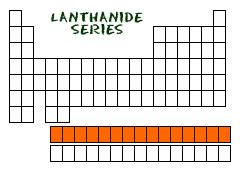
1. They are a metal (duh).



The **Transition Metals** include groups 3 through 12!

Properties: (1) They bond really well to themselves compared to most other elements.

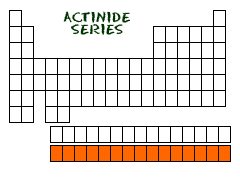
1. They are a metal (duh).



The **Lanthanides** are on the bottom of the periodic table, so we do not give them a group number.

Property: (1) They are shiny metals and similar to alkaline-earth metals.

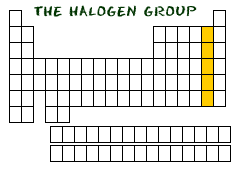
(2) They are also known as rare-earth metals because, well, they are very rare on Earth!



The **Actinides** are on the bottom of the periodic table, so we do not give them a group number.

Properties: (1) They are radioactive because they are unstable and their nucleus will break apart because they are too big.

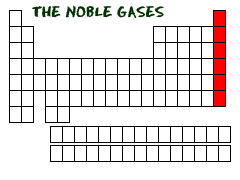
1. Most of these are **synthetic**, meaning they were made in a lab and are not found naturally on Earth.



The **Halogens** are in group 17.

Properties: (1) They are the most reactive non-metal.

(2) They like to react with Alkali Metals (because they “complete” them!)



The **Noble Gases** are group 18.

Properties: (1) They are unreactive because they are “happy” (they have full valence electron orbitals)

(2) They were discovered last of the groups (because of how unreactive they are)