

2.3 Moles, Molecules, and Grams Prequiz

1. What is the Molar Mass of Boron?
 - a) 5 amu
 - b) 10.81 amu
 - c) 10.81 grams/mole
 - d) Period 2
 - e) Group 13
2. What is a Mole in Chemistry?
 - a) The number of electrons in an atom
 - b) A very large number of particles
 - c) The rotational velocity of the electrons in the atom
 - d) The average kinetic velocity of the particles
 - e) A fun dance that Terborg does when he wins a game of Mario Kart
3. How many moles of Carbon are in 12.0 grams of Carbon?
 - a) 12.0 moles
 - b) 1.00 moles
 - c) 144 moles
 - d) 24 moles
 - e) There is not enough information to answer this question
4. What is Avogadro's number?
 - a) 6.022×10^{23}
 - b) 1 million
 - c) 2.998×10^8
 - d) 3.1415...
 - e) Avocado? Now I'm hungry...
5. How many molecules are in 1.00 g of water?
 - a) Exactly 3 molecules
 - b) 3.34×10^{22} molecules
 - c) 5.28×10^{23} molecules
 - d) 7.12×10^{24} molecules
 - e) Now I'm hungry *and* thirsty!
6. What is the molar mass of salt (NaCl)?
 - a) 58.4
 - b) 2
 - c) 13.2
 - d) 21.1
 - e) 8.55
7. When converting between grams, moles, and molecules, which of these three can be directly converted to the other two?
 - a) None of them.
 - b) Grams
 - c) Molecules
 - d) Moles
 - e) All three of them can be directly converted to the other two.

Name _____ Date _____ Period _____

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Answer Key

1. C
2. B
3. B
4. A
5. B
6. A
7. D