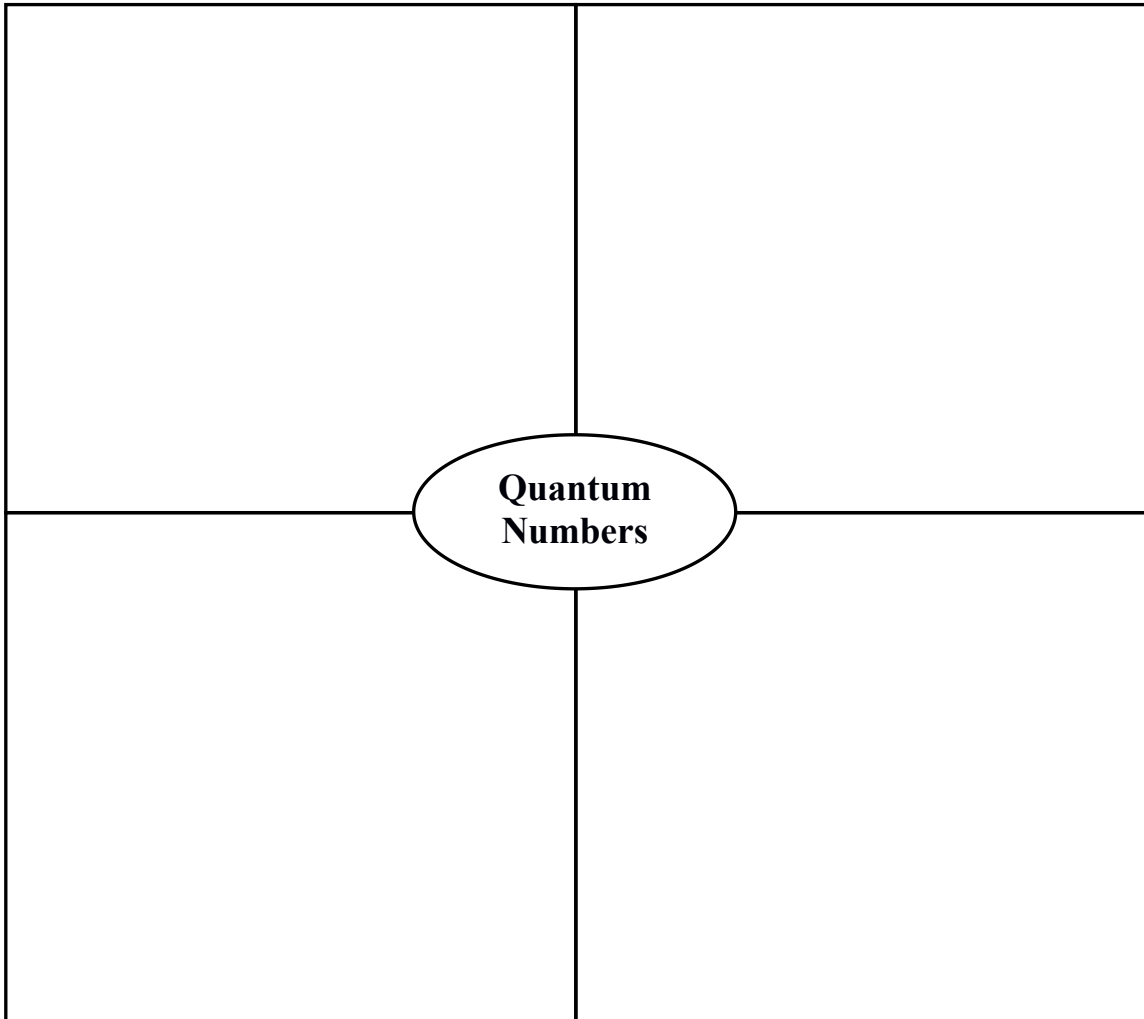


Name _____ Date _____ Period ____

2.2 Orbitals & Electron Configuration Notes

What the atom does NOT look like:

What the atom does look like:



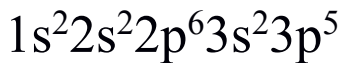
Purpose:

Analogy: Quantum Numbers are like a _____ because

Electron Configuration

Purpose: Writes out the arrangement of all electrons in a single atom.

Energy Level



of Electrons

Orbital Shape

Rules:
Pauli exclusion principle--

Aufbau principle--

Hund's rule--

Examples

Hydrogen -

Helium -

Lithium -

Beryllium -

Boron -

Nitrogen -

Neon -

Sodium -

Aluminum -

Calcium -

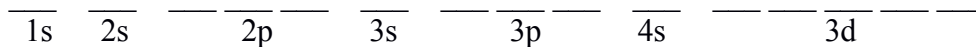
Iron -

Antimony -

Neodymium -

Orbital Diagrams

Iron:



4 Quantum #'s of
this electron?

n=

m=

ℓ=

spin=