**Lewis Diagram Steps**

1. Write the central atom (remember the following tips!)
	1. Carbon is *usually* in the center
	2. Less electronegative elements are *usually* in the center
	3. Hydrogen and Halogens are *usually* at the edge
2. Draw your valence electrons around the center atom.
	1. Remember to use Hund’s rule: one electron on each side before two on any side!
3. Position the remaining atoms so that they match up with unpaired electrons.
4. If it is a polyatomic ion, add or remove electrons and draw brackets around the structure.
5. If there are any unpaired electrons in adjacent atoms, create a double bond.
	1. Look for a possible resonance structure (when there is more than one place for a double bond that results in *different* Lewis Diagrams).
6. Check for the octet rule (every atom wants 8 valence electrons!)
7. Check that the total number of valence electrons matches between the formula & the diagram.

Remember: if you get stuck, perhaps you arranged the atoms incorrectly (step 3). That’s OK, just restart and try a different configuration!

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